Bell Work
26-Jan-2018

Balance the following equation;

$\_Al + \_O_2 \rightarrow \_Al_2O_3$

How many moles **of Aluminum Oxide are produced from the combustion of 3 moles of Aluminum (think about the balanced equation and a molar ratio)**

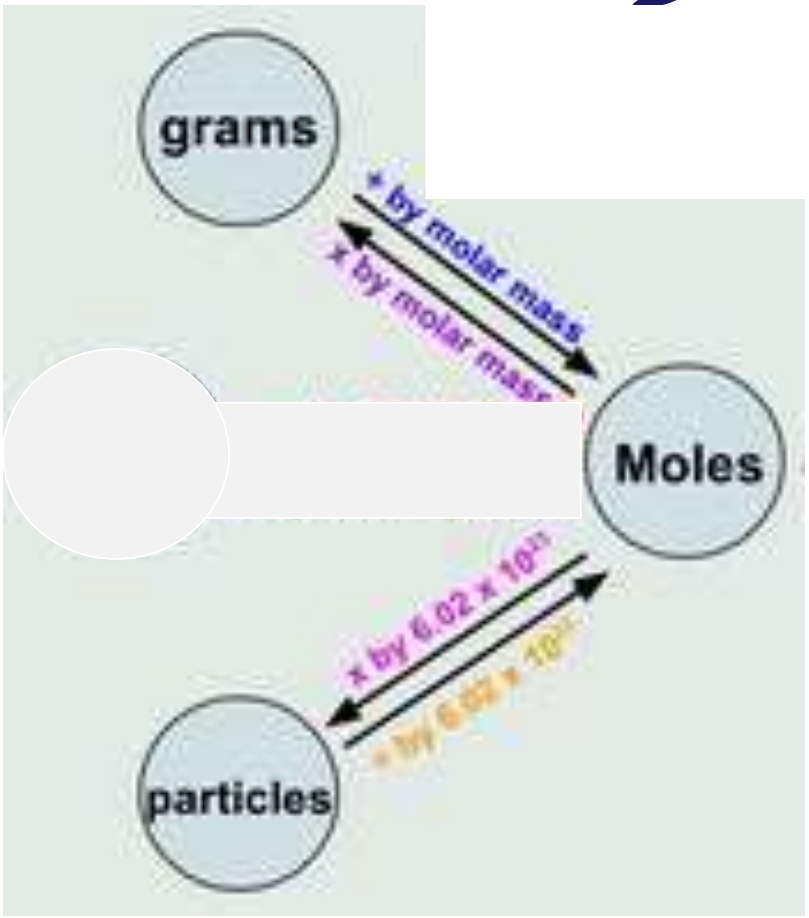
What about grams of O_2 needed?

Objective:

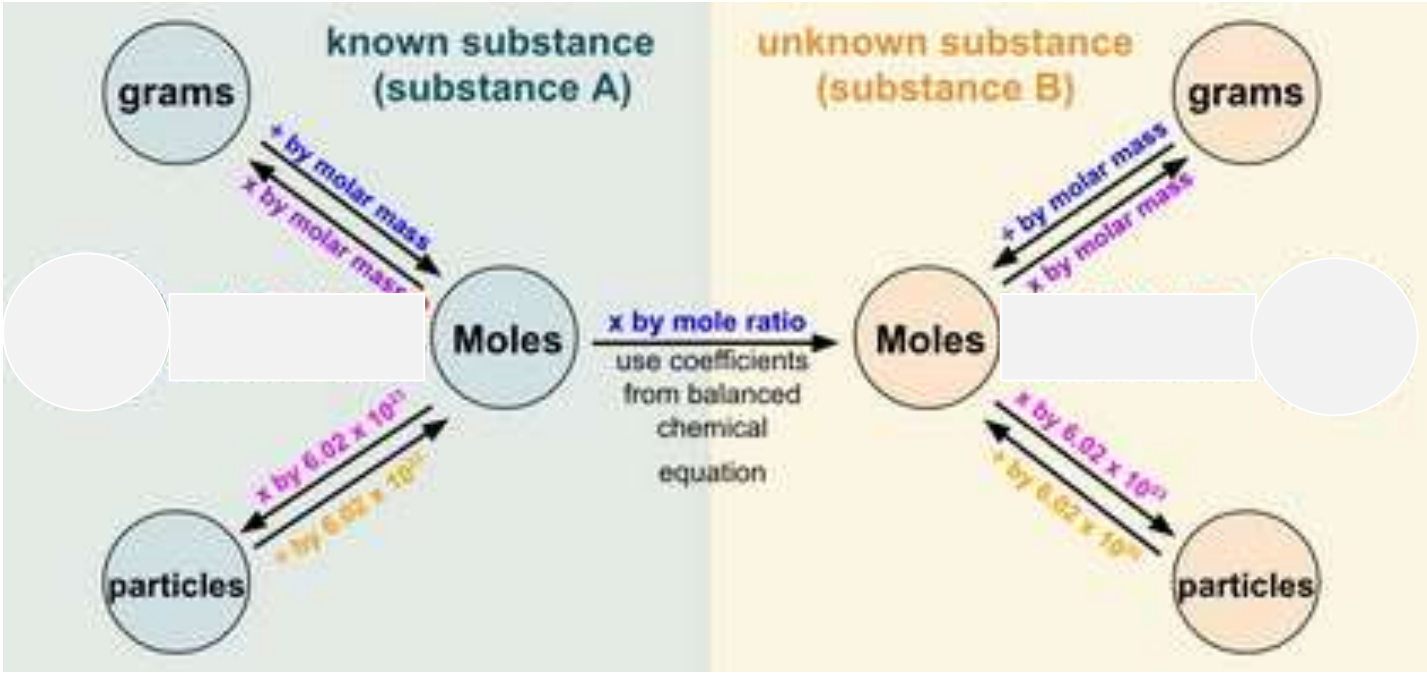
You will be able to set up a mole bridge using a balanced equation.

EQ: After a win on a sports field or an academic assignment what is your next responsibility in order to repeat on the next challenge?

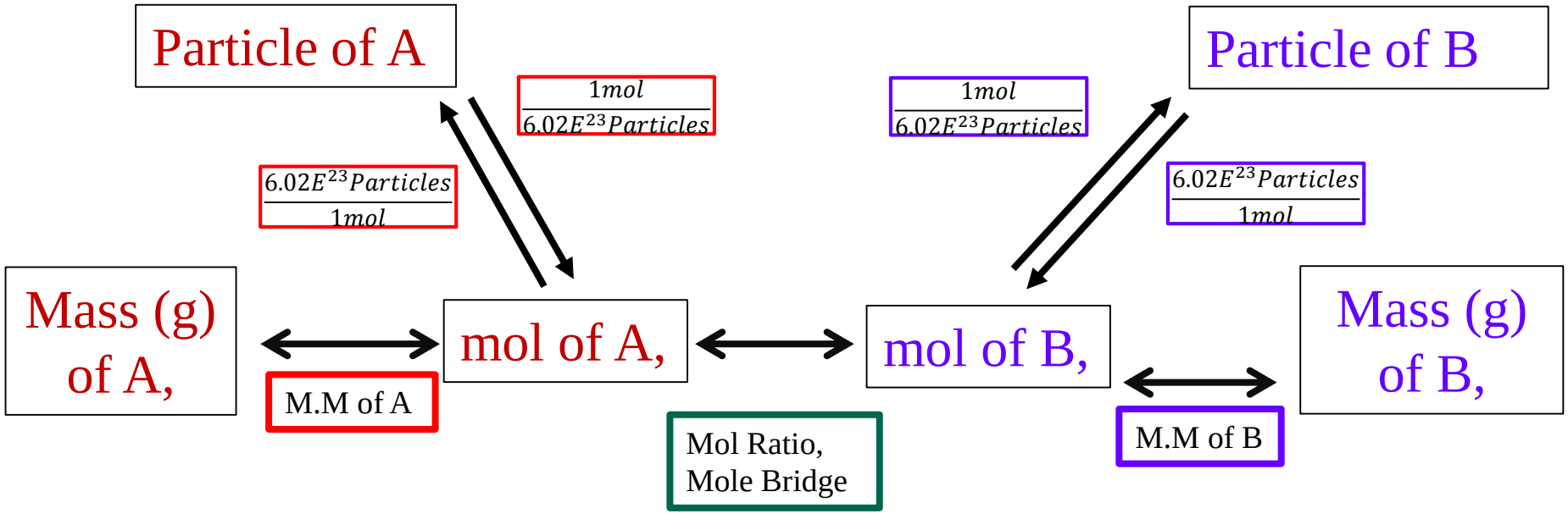
Mole conversion (Same specie)



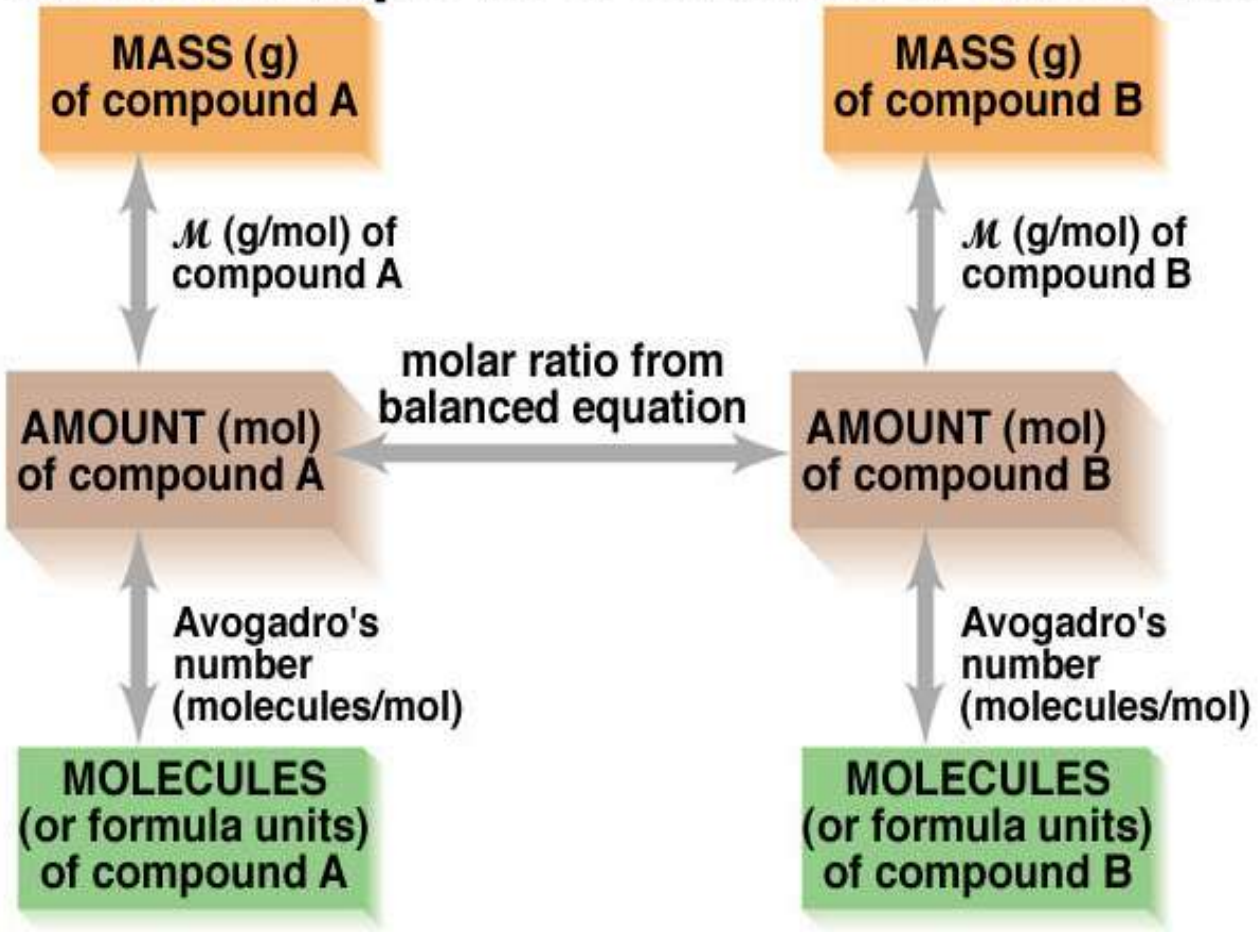
Stoichiometry



Stochiometry Flow Chart



Summary of the Mass-Mole-Number Relationships in a Chemical Reaction



Stoichiometry

Balance the following equation:

$$4\ NH_3 + 5\ O_2 \rightarrow 6\ H_2O + 4\ NO$$

What is the ratio between ammonia and nitrogen monoxide? **4:4 or $\frac{4 mol\ NH_3}{4 mol\ NO}$**

What is the ratio between Nitrogen monoxide and oxygen? **4:5 or $\frac{4\ mol\ NO}{5\ mol\ O_2}$**

So for every 4 mol of NO you have 5 mol of O_2.

Stoichiometry

$$4NH_3 + 5O_2 \rightarrow 6H_2O + 4NO$$

So many conversion factors exist:

4 mol NH_3/5 mol O_2, 6 mol H_2O/4 mol NH_3, etc

What if you had 2 mol of NO, how many moles of O_2 would you have?

$$2\text{mol NO} \times \frac{5\text{mol } O_2}{4\text{mol NO}} = 2.5\text{mol } O_2$$

What if you had 6 mol of H_2O, how many moles of O_2 would you have?

$$6\text{mol } H_2O \times \frac{5\text{mol } O_2}{6\text{mol } H_2O} = 5\text{mol } O_2$$

Stoichiometry

“Stoichiometry” refers to the relative quantities of moles. It also refers to calculations that make use of mole ratios.

Stoichiometry

$$4NH_3 + 5O_2 \rightarrow 6H_2O + 4NO$$

Recall also that molar masses provide factors:

$$\frac{1\ \text{mol } NH_3}{17\ \text{g } NH_3} \qquad \frac{32\ \text{g } O_2}{1\ \text{mol } O_2}$$

Is $\frac{4\ \text{g } NH_3}{5\ \text{g } O_2}$ **a conversion factor?**

No!

The equation tells us moles not grams.

Stoichiometry Question 1a

Consider : $4NH_3 + 5O_2 \rightarrow 6H_2O + 4NO$

How many moles of H_2O are produced if 0.176 mol of O_2 are used?

$$\#\ \text{mol } H_2O = 0.176\ \text{mol } O_2 \times \frac{6\ \text{mol } H_2O}{5\ \text{mol } O_2} =$$

$$0.21\ \text{mol } H_2O$$

Notice: A correctly balanced equation is essential to get the right answer

Stoichiometry questions (1b)

Consider : $4NH_3 + 5O_2 \rightarrow 6H_2O + 4NO$

How many moles of NO are produced in the reaction if 17 mol of H_2O are also produced?

mol NO= $17 \text{ mol } H_2O \times \frac{4 \text{ mol NO}}{6 \text{ mol } H_2O} =$

11.33 mol NO

Notice: A correctly balanced equation is essential to get the right answer!

Recall

What is essential to perform stiochiometry?

Stoichiometry questions 1c

Consider : $4NH_3 + 5O_2 \rightarrow 6H_2O + 4NO$

How many moles of NH_3 are needed in the rxn if 0.5 mol of H_2O are also produced?

$$\#\ \text{mol } NH_3 = 0.5\ \text{mol } H_2O \times \frac{4\ \text{mol } NH_3}{6\ \text{mol } H_2O}$$

= 0.33 mol NH_3

This is what anhydrous ammonia will do to your skin

The "Mole Bridge"

The Mole Bridge is used to convert from one type of compound to another via their molar ratio based on a *Balanced* equation.

You have just converted from one type of compound to another in moles.

Now what if you wanted to go from moles of one compound to grams of another?

The "Mole Bridge"

Now what if you wanted to go from moles of one compound to grams of another?

Moles of A → grams of B

$$\text{mol A} \times \frac{\text{mol B}}{\text{mol A}} \times \frac{\text{Molar Mass B}}{1\ \text{mol B}} = \text{grams B}$$

The Mole Bridge

Your Turn

Consider : $4NH_3 + 5O_2 \rightarrow 6H_2O + 4NO$

How many grams of H_2O are produced if 1.9 mol of NH_3 are combined with excess oxygen?

$$\#\ g\ H_2O = 1.9\ \text{mol } NH_3 \times \frac{6\ \text{mol } H_2O}{4\ \text{mol } NH_3} \times \frac{18\ g\ H_2O}{1\ \text{mol } H_2O} =$$

51 g H_2O

Your Turn

Consider : $4NH_3 + 5O_2 \rightarrow 6H_2O + 4NO$

How many grams of O_2 are required to produce 0.3 mol of H_2O?

$$\#\ g\ O_2 = 0.3\ \text{mol } H_2O \times \frac{5\ \text{mol } O_2}{6\ \text{mol } H_2O} \times \frac{32\ g\ O_2}{1\ \text{mol } O_2} = \mathbf{8\ g\ O_2}$$

Before You Go

Write out the steps for completing stiochiometery.

HW:

Read p. 233-237

Work out all examples and complete #28-32

Bell Work, 2-FEB-2018

Complete bell work in your notes because we will be referencing to it next week for a class lab/ demo.

1. Please write the balanced chemical equation for the reaction of **Iron** and **Copper (II) Sulfate**.

2. Why type of reaction is this?

3. What is the mole:mole ratio of Iron to Copper?

4. Cal. the mass of Iron needed to produce 0.60 g Copper.

Objective – you will be comfortable converting from grams of one substance to grams of another in a balance equation

Moving along the stoichiometry path

Given: $4NH_3 + 5O_2 \rightarrow 6H_2O + 4NO$

a) How many moles of H_2O can be made using 0.5 mol NH_3?

b) What mass of NH_3 is needed to make 1.5 mol NO?

c) How many grams of NO can be made from 120 g of NH_3?

Converting grams ←→ grams

Notice that we cannot directly convert from grams of one compound to grams of another. Instead we have to go through moles.

Many stoichiometry problems follow a pattern:

gram(A) ↔ mol(A) ↔ mol(B) ↔ gram(B)

grams of A → grams of B

Converting grams ↔ grams

Many stoichiometry problems follow a pattern:

gram(A) ↔ mol(A) ↔ mol(B) ↔ gram(B)

grams of A → grams of B

$$\text{gram A} \times \frac{1\ \text{mol A}}{\text{MM A}} \times \frac{\text{mol B}}{\text{mol A}} \times \frac{\text{MM B}}{1\ \text{mol B}} = \text{grams B}$$

Grams A to Moles A

Mole Bridge

Moles B to Grams B

Your Turn… You need to think a little harder!

Consider : $4NH_3 + 5O_2 \rightarrow 6H_2O + 4NO$

How many grams of NO is produced if 12g of O_2 is combined with excess ammonia?

\# g NO=

$$12\ \cancel{g\ O_2} \times \frac{1\ \cancel{mol\ O_2}}{32\ \cancel{g\ O_2}} \times \frac{4\ \cancel{mol\ NO}}{5\ \cancel{mol\ O_2}} \times \frac{30\ g\ NO}{1\ \cancel{mol\ NO}}$$

$$= \mathbf{9.0\ g\ NO}$$

Converting grams to grams

grams(A) $\leftrightarrow$ moles(A) $\leftrightarrow$ moles(B) $\leftrightarrow$ grams(B)
We can start anywhere along this path

So, for the rxn $2H_2 + O_2 \rightarrow 2H_2O$ what is the path we would take for the following

Given 2 mol H_2O, calculate grams H_2O?
Moles O_2 required for 36 g H_2?
Grams of H_2O produced from 6g O_2?

36g **9moles** **6.75g**

Answers

$$4NH_3 + 5O_2 \rightarrow 6H_2O + 4NO$$

a)

$$\#\ \text{mol } H_2O = 0.5\ \text{mol } NH_3 \times \frac{6\ \text{mol } H_2O}{4\ \text{mol } NH_3} = 0.75\ \text{mol } H_2O$$

b)

$$\#\ \text{g } NH_3 = 1.5\ \text{mol NO} \times \frac{4\ \text{mol } NH_3}{4\ \text{mol NO}} \times \frac{17.04\ \text{g } NH_3}{1\ \text{mol } NH_3} = 25.6\ \text{g } NH_3$$

c)

$$\#\ \text{g NO} = 120\ \text{g } NH_3 \times \frac{1\ \text{mol } NH_3}{17.04\ \text{g } NH_3} \times \frac{4\ \text{mol NO}}{4\ \text{mol } NH_3} \times \frac{30.01\ \text{g NO}}{1\ \text{mol NO}}$$

$$= 211\ \text{g NO}$$

Small Group

Stiochiometry Practice #2,

You will complete Stiochiometry Practice #2 in a small group.

As a team, work out each problem on white boards, showing all work.

More Stoichiometry Questions

1. **$2\ C_4H_{10} + 13\ O_2 \rightarrow 8\ CO_2 + 10\ H_2O$**

a) What mass of O_2 will react with 400 g C_4H_{10}?

b) How many moles of water are formed in a)?

a)1434g b)34.5moles

2. **$3\ HCl + Al(OH)_3 \rightarrow 3\ H_2O + AlCl_3$**

How many grams of aluminum hydroxide will react with 5.3 moles of HCl? 137.8g

3. **$Ca(ClO_3)_2 \rightarrow CaCl_2 + 3\ O_2$**

What mass of O_2 results from the decomposition of 1.00 kg of calcium chlorate? 466g

Continued

4. $Fe_2O_3 + 3CO \rightarrow 2Fe + 3CO_2$. a) 3.09g b) 28.3g

A) How many moles of carbon monoxide are required to react with 163.0 g of iron(III) oxide?

B) How many grams of CO_2 are produced from a reaction that also produces 23.9 grams of Fe?

5. $3Cu + 8HNO_3 \rightarrow 3Cu(NO_3)_2 + 4H_2O + 2NO$

A) How many moles of copper(II) nitrate can be prepared from 17.0 moles of Cu?

B) How many grams of copper(II) nitrate can be prepared using 3.8 moles of HNO_3?

C) What mass of water results from the reaction of 8.50 kg of copper metal?

Bell Work
6-Feb-2018

1. Name the following compounds:

$(NH_4)_2Cr_2O_7$ **PCl_5** **$Mo(PO_4)_2$**

2. What are the molar masses of each compound?
3. What do the coeficients infront of species in a balance equation represent?

Agenda

Percent Yeild

Finish Lab, Day 1

Objective:

You understand that some reactants run out before others and will be able to determine the limiting reagent in a reaction

Percent Yield

Rxns rarely produce the predicted amount of product from the masses of reactants in the rxn.

$1Fe(s) + 1CuSO_4(aq) \rightarrow 1FeSO_4(aq) + 1Cu\ (s)$

An example of this is the rxn of $CuSO_4$ with Fe. Normally we expect a 1 mol yield of Cu for every mol of Fe reacted. This does not always happen.

Percent yield

If you react 55.8 g of Fe to make Cu, the amount of Cu expected is 1 mol of Cu or 63.5 g of Cu.

$1Fe(s) + 1CuSO_4(aq) \rightarrow 1FeSO_4(aq) + 1Cu(s)$

Sadly the amount you get will probably be < 63.5g say, 50 g of Cu. The problem is a competing rxn or complexing of a Cu^{2+} ion that happens.

Percent Yield

The Cu^{2+} participating in this "complexing" will not be able to make Cu. The reaction will not yield 100% of the expected Cu.

The amount of Cu produced, ~50 g is only 78.7% and not 100 % of the expected 63.5 g.

Percent yield = 100 x $\frac{\text{50 g Cu actual}}{\text{63.5 g Cu predicted}}$ → 78.8%

Percent Yield

$$\text{Percent yield} = \frac{\text{Actual yield}}{\text{Theoretical yield}} \times 100$$

What is the percent yield for a rxn if you predicted the formation of 21 g of C_6H_{12} and actually recovered only 3.8 g?

Percent yield = 100 x $\frac{\text{3.8 g } C_6H_{12} \text{ actual}}{\text{21 g } C_6H_{12} \text{ predicted}}$ → 18%

Percent Yield Demo

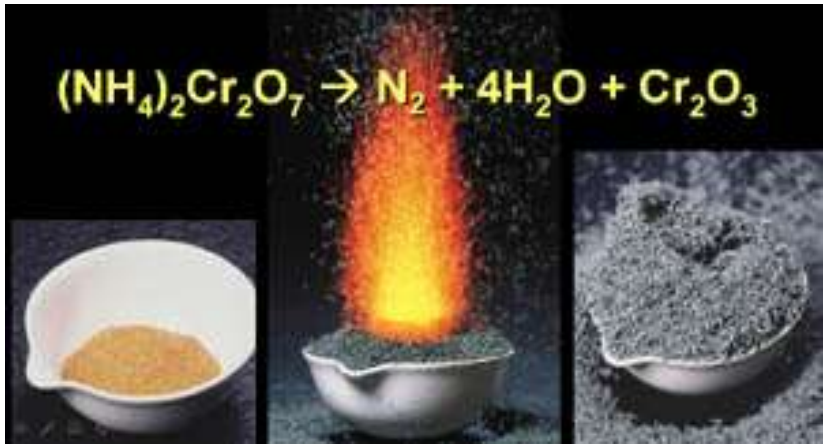
When the following rxn is called a mini "volcano"

$(NH_4)_2Cr_2O_7 \rightarrow N_2 + H_2O + Cr_2O_3$

Balance it

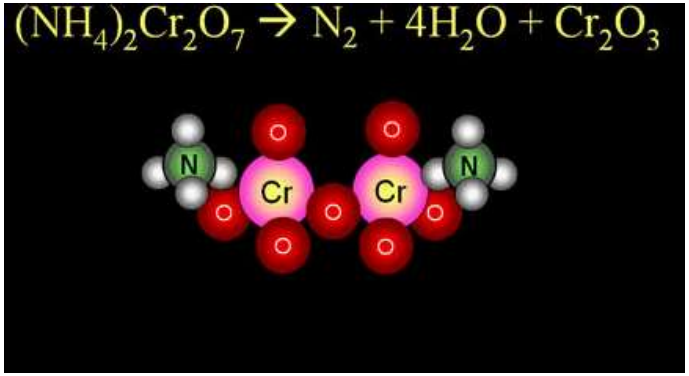
What type of reaction is this?

When 3.0g of the reactant is heated how many grams of the chromium containing product would you expect to get?



Percent Yield Demo

When the following reaction takes place it is called a mini "volcano"



$$\% \text{ Yield} = \frac{\text{Actual}}{\text{Theoretical}} \times 100$$

If we actually got 1.1g of chromium (III) oxide, what is our percent yield?

Bell Work
7-Feb-2018

On a blank piece of paper… Start a prelab for the "S'mores Lab" – after Purpose and Safety being prelab questions.

This lab is your lesson for the next part of the Stoichiometry Unit; *Limiting Reagents*.

All work must be shown in accordance to how Mr. Golden models every example problem. Failure to do so will earn a zero "0" and Mr. Golden will be unable to help you.

Agenda

S'mores Limiting Reagent Stiochiometry Lab

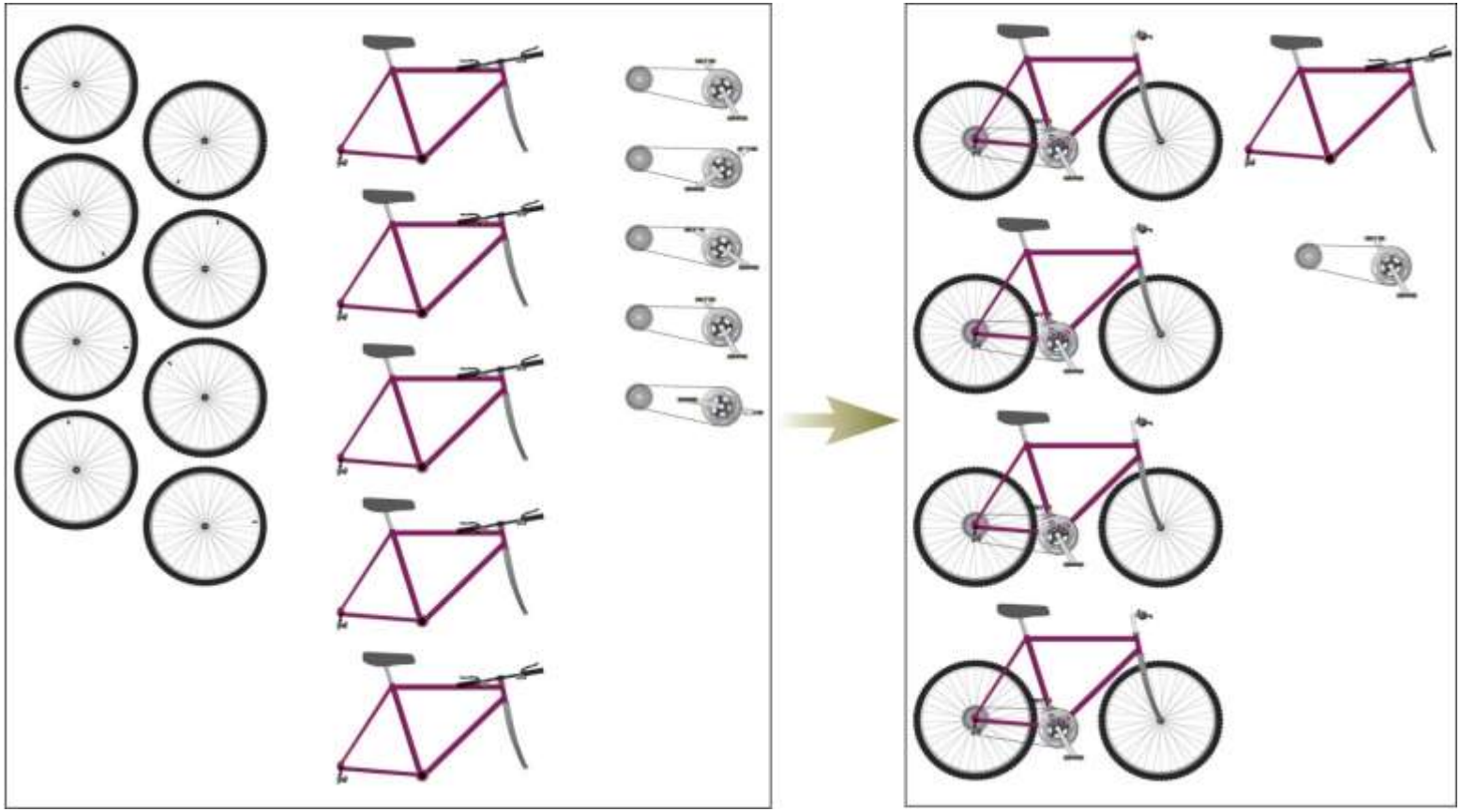
Objective

You will be able to define what a limiting reagent is and how it is relevant in stoichiometry.

You will be able to identify if a chemistry stoichiometry problem requires a limiting reagent calculation.

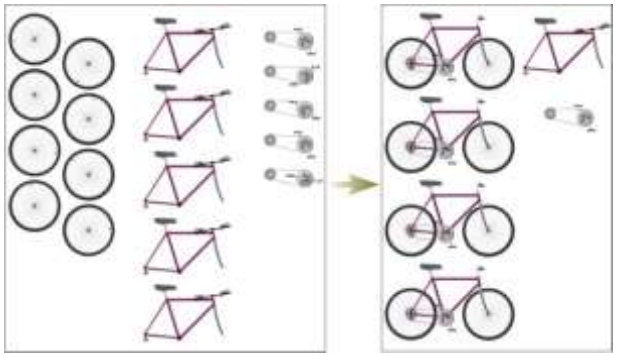
Limiting Reagent

The reactant that is completely consumed by the reaction.



Limiting Reagent

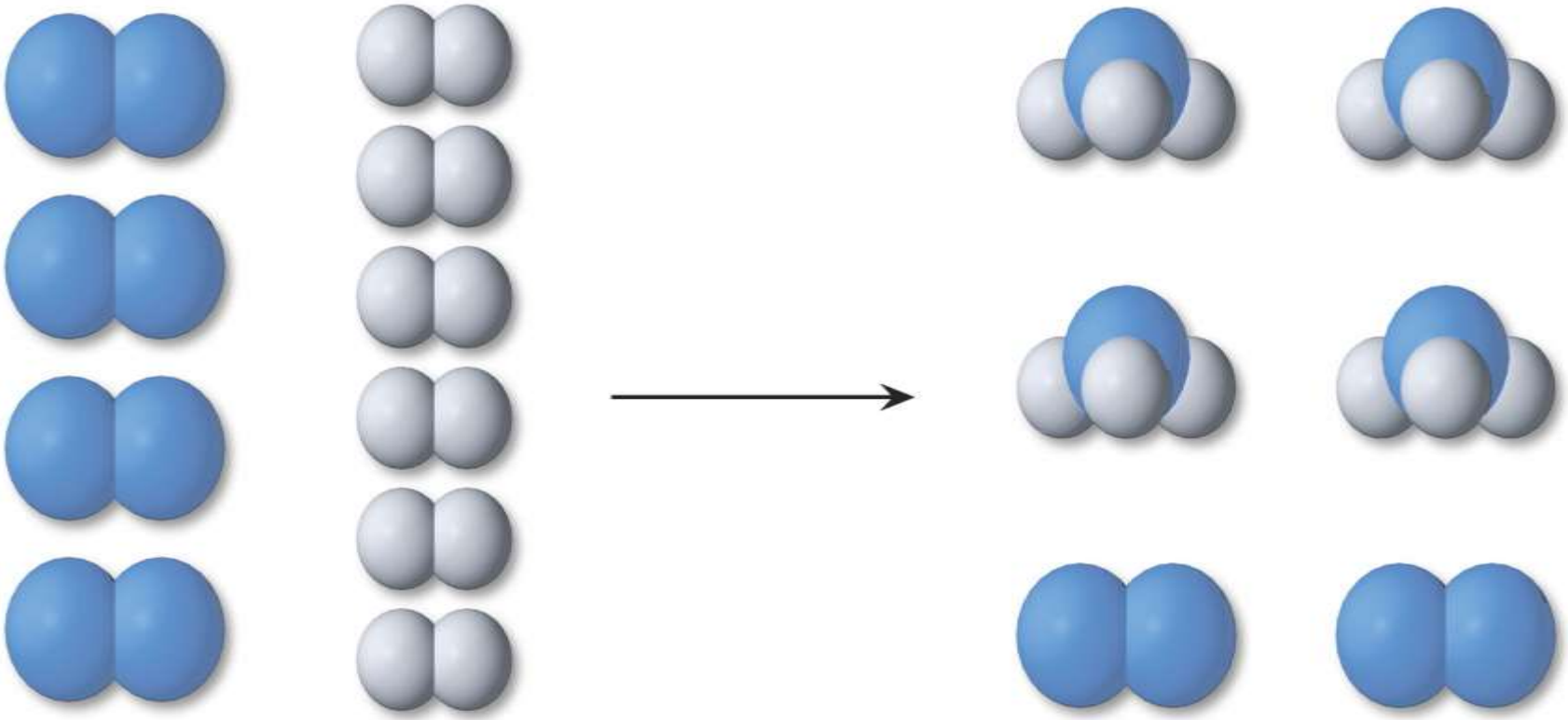
The reactant that is completely consumed by the reaction



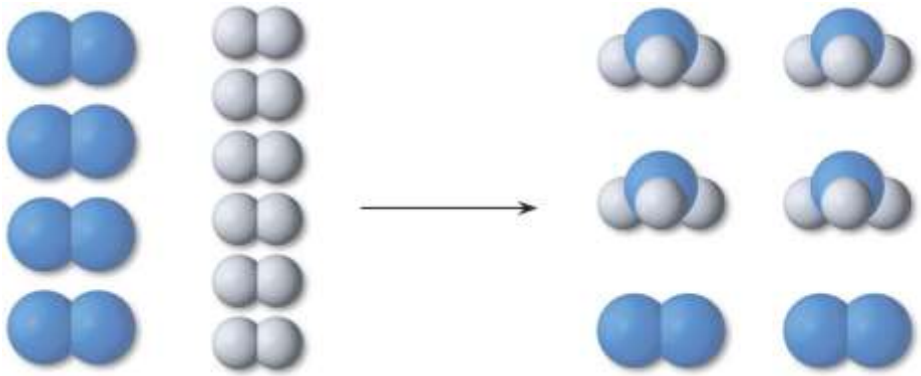
The number of bicycles that can be assembled is limited by whichever part runs out first. In the inventory shown in this figure, wheels are that part.

Limiting Reagent

$N_2(g) + 3H_2(g) \rightarrow 2NH_3(g)$



Limiting Reagent



A molecular view of a Limiting reactant situation for the ammonia synthesis. To make 4 molecules of NH_3 requires 2 molecules of N_2 & 6 molecules of H_2. If we start with 4 molecules of N_2 and 6 molecules of H_2, H_2 is the limiting reactant.

Limiting Reagent

So to find limiting reagents carry out the stiochiometry for each of your reactants that you are give a quantity for.
The reactant that gives you the smallest value of product is your limiting reagent.
Use this reactant for all calculations

Bell Work
8-Feb-2018

a. If lead (II) nitrate reacts with potassium iodide what are the two products?

b. What type of reaction is this?

c. Write out a balanced net ionic equation.

d. If you have 2moles of lead (II) nitrate how many grams of potassium iodide would you need for a complete rxn?

Agenda

S'mores limiting reagent stiochiometry lab and practice

Objective

You will be able to determine the limiting reagent (LR) and non limiting reagent (NLR) in a reaction in stoichiometry.

You will be able to solve stoichiometric relationships according to the LR

Limiting Reagent

$$4KO_{2(s)} + 2H_2O_{(l)} \rightarrow 4KOH_{(s)} + 3O_{2(g)}$$

a. How many moles of O_2 can be produced from 0.15 mol KO_2 and 0.10 mol H_2O?

b. Determine the limiting reactant.

$$4KO_2(s) + 2H_2O(l) \rightarrow 4KOH(s) + 3O_2(g)$$

0.15 mol **0.10 mol** **? moles**

Limiting Reagent

b. Determine the limiting reactant.

$4KO_2(s) + 2H_2O(l) \rightarrow 4KOH(s) + 3O_2(g)$

0.15 mol 0.10 mol ? moles

Based on **KO_2** :

0.15 mol ~~**KO_2** x~~ $\frac{3\text{mol } \mathbf{O_2}}{4\text{mol } \mathbf{KO_2}}$ **= 0.1125 mol O_2**

Limiting Reagent

b. Determine the limiting reactant.

$4KO_2(s) + 2H_2O(l) \rightarrow 4KOH(s) + 3O_2(g)$

0.15 mol 0.10 mol ? moles

Based on **H_2O:**

$$0.10 \text{ mol } \mathbf{H_2O} \times \frac{3 \text{mol } \mathbf{O_2}}{2 \text{mol } \mathbf{H_2O}} = \mathbf{0.15 \text{ mol } O_2}$$

Limiting Reagent

b. Determine the limiting reactant.

$4KO_2(s) + 2H_2O(l) \rightarrow 4KOH(s) + 3O_2(g)$

0.15 mol 0.10 mol ? moles

Based on **KO_2** : **= 0.1125 mol O_2**

Based on **H_2O:** **= 0.15 mol O_2**

KO_2 is the limiting reagent because it limited the amount of O_2 that could be produced. H_2O is the excess reagent.

Practice questions

1. **$2Al + 6HCl \rightarrow 2AlCl_3 + 3H_2$**
   If 25 g of aluminum was added to 90 g of HCl, what mass of H_2 will be produced?

2. **$N_2 + 3H_2 \rightarrow 2NH_3$: If you have 20 g of N_2 and 5.0 g of H_2, which is the limiting reagent?**

Practice questions

2. **$N_2 + 3H_2 \rightarrow 2NH_3$: If you have 20 g of N_2 and 5.0 g of H_2, which is the limiting reagent?**

3. **What mass of aluminum oxide is formed when 10.0 g of Al is burned in 20.0 g of O_2?**

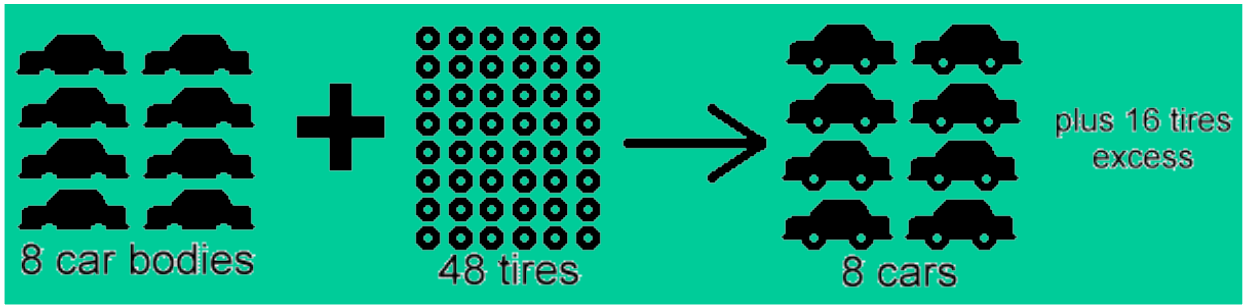
S'mores Lab

You will be making one these in the lab…

…and solving post lab questions #1-17.

S'mores Lab

S'mores are just like stiochiometry and chemistry you are limited to how many you can make, through this we briefly look at the concept of limiting reagent,



S'mores Lab

You will use one (1) gram cracker, which you will split in half

ONE (nugget) of chocolate

One marshmallow